

Topic 1 Stoichiometric Relationships

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Stoichiometry: the quantitative method of examining the relative amounts of reactant and products. Limiting agent: the reactant that will be completely consumed during the reaction. Yields. Theoretical yield: the yield that is calculated. Experimental yield: the yield that is obtained. Difference between yields due to: impurities

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Today we will cover Topic 1: Stoichiometric Relationships This is one of the big 3 topics (the others being Bonding and Organic Chemistry). These 3 topics make up almost 50% of the marks on papers 1 and 2. One needs to practice and be ready to convert volumes of known solutions, masses of solids and volumes of gases into moles.

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Topic 1.1 Introduction to the particulate nature of matter Atoms of different elements combine in fixed ratios to form compounds, which have different properties from their component elements. Mixtures contain more than one element and/or compound that are not chemically bonded together and so retain their individual properties.

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Notes for the Core IB Chemistry module: Topic 1: Stoichiometric relationships. These have been made according to the specification and cover all the relevant topics for examination in May/June. Core IB Topic 1: Stoichiometric relationships Notes: Chemical Equations; Formulas; Mass and Gaseous Volume Relationships in Chemical Reactions; Solutions; The Mole

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Topic 1: Stoichiometric relationships 13.5 hours

Task: 'Topic 1: Stoichiometric relationships' contains many scientific laws. One of the most important of these is the Ideal Gas Law - $PV=nRT$. Using this as an example, which of the features below might be able to categorise a law in chemistry?

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Topic 1: Stoichiometric relationships. 1.2 Formulas · The molecular formula shows the number of atoms of each element present in a molecule. · The empirical formula gives the ratio of the atoms of different elements in a compound. It is the molecular formula expressed as its simplest ratio.

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Topic 1. Stoichiometric Relationships. STUDY. PLAY. atomic theory. 1. All matter is composed of atoms. 2. Matter cannot be created or destroyed, just rearranged during chemical reactions. 3. Physical and chemical properties of matter depend on bonding and the arrangement of these atoms.

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