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~~Heat~~

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The difference between thermal energy and heat is that thermal energy is not being transferred, but remains part of the system's internal energy (kinetic

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The Laws of

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expansion, specific heat, and heats of fusion and vaporization.

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Why? Macroscopic, because atoms were unknown at the time thermodynamics was created.

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The second law of thermodynamics. Heat does not flow spontaneously from a colder region to a hotter region, or, equivalently, heat at a given temperature cannot be converted entirely

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Definition of temperature. The

definition of thermodynamic temperature T is a function of the change in the system's entropy S under reversible heat transfer Q rev: =. Entropy being a state function, the integral of dS over any cyclical process is zero. For a system in which the entropy is purely a function of the system's energy E, the temperature can be defined as:

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