

Nelson Chemistry 12 Chapter 5 Review Answers

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FSc Chemistry Book2, CH 5, LEC 6: Disproportionation Reactions of Halogens

Zumdahl Chemistry 7th ed. Chapter 5 (Part 1) **Chemistry 12th | Chapter 5 Surface chemistry(?????? ??????) | NCERT Based All topics in one video**

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 $q = mc\Delta T = (10.5 \text{ }^\circ\text{C}) (1.00 \text{ g}) (4.18 \text{ J/g}^\circ\text{C}) = 43.9 \text{ J}$ Statement:
The amount of energy absorbed by 1.00 g of water, $\text{H}_2\text{O(l)}$, when its temperature is raised $10.5 \text{ }^\circ\text{C}$ is 43.9 J. 36.
Given: $V \text{ HCl(aq)} = 100.0 \text{ mL}$; $V \text{ NaOH(aq)} = 100.0 \text{ mL}$; $T_{\text{initial}} = 23.5 \text{ }^\circ\text{C}$; $T_{\text{final}} = 36.0 \text{ }^\circ\text{C}$

Chapter 5 Review, pages 338–339 - Pre University Courses

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Section 6.5: Rate Law. Tutorial 1 Practice, page 380 1.

(a) The order of reaction with respect to A is 2. When the initial concentration of A is doubled, the rate increases by a factor of 4, or 2^2 , so, the exponent of the concentration of reactant A is 2. The order of reaction with respect to B is 1. When the initial concentration of B is doubled, the rate is doubled, or multiplied by 2^1 , so the exponent of the concentration of reactant B is 1.

Section 6.5: Rate Law - Pre University Courses

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