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Revision Chemistry \"Reacting Masses 1\
Equilibrium 1*

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\ "Calculating Moles of an Element\"

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University Affordable Learning Solutions Program, and Merlot. We also acknowledge previous National Science Foundation support under grant numbers 1246120, 1525057, and 1413739.

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The infra-red spectra of H, I and J are given below, but it does indicate which is - which. Identify the three compounds H, I and J, using the infra-red spectra below, and decide which spectrum belongs to which compound.

(iii) Compound E, which is a branched chain

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haloalkane, was found to have the composition by mass of 39.8% C, 7.3% H, 2)

IR TASK 1 - Weebly

1) 2) 3) 4) 5) 6) bromoethane - ethyl ethanoate cold, aq NaOH nucleophilic substitution propene 2-methylpropylamine chloroethane ethanoic acid

ORGANIC SYNTHESIS PROBLEMS

2-bromo-2-methylbutane 4 5-40 (CH₃ CH₂), 5-40 (CH₃ CH₂), 10-70 (CBr), 5-40 ((CH₃)₂ CBr) methylpropene) 3 90-150 (CH₂), 90 -150 (C₅ 40 H 3 2) propene 3) 90-150 (CH₂), 90

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-150 (C₅ H₁₀ Cl) 2-chloropropane), 2 5-40 (CH₃ C=O) propanone), 2 20-50 (CH₃ C=O) 190-220 (CO) methylamine 1 25-60 (CH₃)

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2,2-dimethylamine 2-methylbut-1-ene

3-methylbutan-1-ol 4-hydroxybutanone

3-methylbutanoic acid butyl methanoate

ethanenitrile propanal Decide which spectrum belongs to which compound and draw the molecule next to the spectrum.

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infra-red spectra of H, I and J are given below, but it does indicate which is - which. Identify the three compounds H, I and J, using the infra-red spectra below, Page 4/25

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Structure	Spectrum	Bond(s)	Wavenumber range (cm ⁻¹)
butanoic ...			

IR PROBLEMS 1 The IR spectra of six compounds are shown ...

rate = k[A][B]² k = rate = 16 = 2.0 mol⁻² dm⁶ s⁻¹ [A][B]² (2 x 2²) b) rate = k[E] k = rate = 0.80 = 0.20 s⁻¹ [E]⁴ c) rate = k[P][Q] k =

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rate = 0.00200 = 0.20 mol⁻¹ dm³ s⁻¹ [P][Q]
0.1 x 0.1 d) rate = k[H⁺][CH₃COCH₃] k =
rate = 9 x 10⁻⁵ = 3.0 x 10⁻⁵ mol⁻¹ dm³ s⁻¹
[H⁺][CH₃COCH₃] (0.5 x 6.0)

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1) 2) 3) 4) 5) 6) bromoethane - ethyl
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2,2-dimethylamine 2-methylbut-1-ene

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3-methylbutanoic acid butyl methanoate

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Identify the three compounds H, I and J, using the infra-red spectra below, Page 4/25

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spectra to the compounds. Identify any key signals you used to identify each compound. You may not be able to decide between two of the compounds. Structure Spectrum Bond(s) Wavenumber range (cm⁻¹) Structure Spectrum Bond(s) Wavenumber range (cm⁻¹) butanoic ...

IR PROBLEMS 1 The IR spectra of six compounds are shown ...

rate = $k[A][B]^2$ $k = \frac{\text{rate}}{[A][B]^2} = \frac{16}{2.0 \times 2.0^2} = 2.0 \text{ mol}^{-2} \text{ dm}^6 \text{ s}^{-1}$
a) rate = $k[E]^4$ $k = \frac{\text{rate}}{[E]^4} = \frac{0.80}{0.20^4} = 0.20 \text{ s}^{-1}$
b) rate = $k[P][Q]$ $k = \frac{\text{rate}}{[P][Q]} = \frac{0.00200}{0.1 \times 0.1} = 0.20 \text{ mol}^{-1} \text{ dm}^3 \text{ s}^{-1}$
c) rate = $k[H^+][CH_3COCH_3]$ $k =$

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$$\text{rate} = 9 \times 10^{-5} = 3.0 \times 10^{-5} \text{ mol}^{-1} \text{ dm}^3 \text{ s}^{-1}$$
$$[\text{H}^+][\text{CH}_3\text{COCH}_3] (0.5 \times 6.0)$$

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