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textbook, read about synthesis, combustion, decomposition, and replacement reactions. Assume that Q, T, X, and Z are symbols for elements. Match each equation in Column A with the reaction type it

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represents in Column B.

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reactions, condensation reactions (and the reverse, cleavage reactions), and oxidation–reduction reactions. The general forms of these five kinds of reactions are

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summarized in Table, along
with examples of each.

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following chemical reactions, write a word equation, a skeleton equation, and a balanced chemical equation. Be sure to show the state of each reactant and product. If you need more help writing

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formulas or determining the state of a substance, refer to Chapters 7 and 8 and the periodic table on pages 178—179. 17.

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Know some of the factors that
cause a chemical reaction to
occur. 3. Be able to identify the
solid that forms in a
precipitation reaction. 4. Be

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able to write and balance molecular, complete ionic, and net ionic equations for a chemical reaction. 5. Be able to write a balanced chemical equation for the reaction between an acid and a base. 6.

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synthesis, combustion, decomposition, and replacement reactions. Assume that Q, T, X, and Z are symbols for elements. Match each equation in Column A with the reaction type it represents in

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Column B. Section 10.2

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Not all single replacement
reactions actually occur.

Reactivity of elements
determines whether or not the
reaction takes place. If the

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reactivity of the single element is greater than the reactivity of the element in the compound it is attempting to replace - the reaction will happen.

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 $C (s) + O_2 (g) \rightarrow CO_2 (g)$
 $H_2O (l) + SO_3 (g) \rightarrow H_2SO_4$
(aq) Decomposition reactions.
A single reactant breaks down
to form 2 or more products.

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Examples: $\text{H}_2\text{CO}_3(\text{aq}) \rightarrow \text{H}_2\text{O}(\text{l}) + \text{CO}_2(\text{g})$
 $\text{CaCO}_3(\text{s}) \rightarrow \text{CaO}(\text{s}) + \text{CO}_2(\text{g})$ Single-replacement reactions. A single element replaces a similar element of an adjacent reactant compound.

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with the reaction type it
represents in Column B.
Column A Column B 1.

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Chemical reactions occur when
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bonds break between the reactants, and atoms rearrange to form new chemical bonds. These reconfigured atoms can either be a compound or just an element. The products have

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molecules are converted to the products. In some cases, reactions can also be ...

Classify the following reaction:
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Most chemical reactions can be classified into one or more of five basic types: acid-base reactions, exchange reactions, condensation reactions (and the reverse, cleavage reactions), and

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oxidation-reduction reactions.
The general forms of these five
kinds of reactions are
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with the reaction type it represents in Column B.

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condensation reactions (and
the reverse, cleavage
reactions), and
oxidation–reduction reactions.
The general forms of these five

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kinds of reactions are summarized in Table, along with examples of each.

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you need more help writing formulas or determining the state of a substance, refer to Chapters 7 and 8 and the periodic table on pages 178—179. 17.

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Know some of the factors that
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solid that forms in a

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precipitation reaction. 4. Be able to write and balance molecular, complete ionic, and net ionic equations for a chemical reaction. 5. Be able to write a balanced chemical equation for the reaction

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reaction type it represents in
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reaction takes place. If the reactivity of the single element is greater than the reactivity of the element in the compound it is attempting to replace - the reaction will happen.

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 $H_2O(l) + SO_3(g) \rightarrow H_2SO_4(aq)$
Decomposition reactions.
A single reactant breaks down

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to form 2 or more products.

Examples: $\text{H}_2\text{CO}_3(\text{aq}) \rightarrow \text{H}_2\text{O}(\text{l}) + \text{CO}_2(\text{g})$

$\text{CaCO}_3(\text{s}) \rightarrow \text{CaO}(\text{s}) + \text{CO}_2(\text{g})$

Single-replacement reactions. A single element replaces a similar element of an adjacent

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each equation in Column A
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