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103 Chemistry I Laboratory
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2: Determination of the
Molarity of a Sulfuric
Acid Solution strans Trial
1 Trial 2 Trial 3 Trial 4*

Starting Buret Reading:
|•74 L 0n15L 0-59L Ending
Buret Reading: 1l.83
L10-10L 10.591 mL of

Titrant Used:

10.09L/9.95L10.00L mmole
of NaOH: 1.009 mol 9.95
no 1.0 ml mmole of H₂SO₄
0.5 mal 5.0 0.5 5.0nd

Molarity of H₂SO₄ Solution
0.05 M 0.5 M 0.05 M 0.05M

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Average Molarity 0.6M.

*Answered: 103 Chemistry I
Laboratory Manual, 2019... |
bartleby*

*Grades: The lab component
counts as 25% of your
final grade in CHEM 103.
Per department policy, you
must perform every
experiment in order to
pass CHEM 103. Grades are
assigned on a numerical
basis out of 220 possible
points. Grade breakdown is
shown below: Lab
etiquette: 12 points (1
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chemistry laboratory and
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Reaction 1: $1 \text{ NaHCO}_3 (s)$
 $1 \text{ NaOH} (s) + 1 \text{ CO}_2 (g)$

Reaction 2: $2 \text{ NaHCO}_3 (s)$
 $1 \text{ Na}_2\text{CO}_3 (s) + 1 \text{ CO}_2$

$(g) + 1 \text{ H}_2\text{O} (g)$ Reaction

3: $2 \text{ NaHCO}_3 (s) \rightarrow 1 \text{ Na}_2\text{O}$

$(s) + 2 \text{ H}_2\text{O} (l)$

Propose a voltaic cell that would coat a piece of copper with platinum. Include the half reactions for your cell and the voltage potential based on the.

CHEM 103 - UIUC

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Laboratory Manual, Fall

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in room 1375 (chemistry computer lab) during the first two weeks of classes and later through the first-floor stockroom. Lab notebooks (with carbonless copies) purchased from the University Bookstore are also acceptable.

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in the stockroom 111. Band-
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answer is your local
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(Example: A K^+ salt, a halogen, and an acid.)

2. Identify the specific compound from each group.

(Example: KCl , Br_2 and HCl)

3. Explain the reasoning that enabled you to come to these conclusions.

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Buret Reading: 11.83
L 10-10L 10.591 mL of
Titrant Used:
10.09L/9.95L 10.00L mmole
of NaOH: 1.009 mol 9.95
no 1.0 mL mmole of H₂SO₄
0.5 mol 5.0 0.5 5.0nd
Molarity of H₂SO₄ Solution
0.05 M 0.5 M 0.05 M 0.05M
Average Molarity 0.6M.

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Grades: The lab component
counts as 25% of your
final grade in CHEM 103.
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 $(s) + 2 \text{ CO}_2 (g) + \text{H}_2\text{O} (g)$
Propose a voltaic
cell that would coat a
piece of copper with
platinum. Include the half

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reactions for your cell and the voltage potential based on the.

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