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9.1. Objectives. •Define stoichiometry. •Describe the importance of the mole ratio in stoichiometric calculations. •Write a mole ratio relating two substances in a chemical equation. 2/26/2018 2. Stoich Definitions. •Composition stoichiometry deals with the mass relationships of elements in compounds. •Reaction stoichiometry involves the mass relationships between reactants and products in a chemical reaction.

Chapter 9

stoichiometry (which you studied in Chapter 3) deals with the mass relationships of elements in compounds. Reaction stoichiometry involves the mass relationships between reactants and products in a chemical reaction. Reaction stoichiometry is the subject of this chapter and it is based on

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Stoichiometry Chapter 9, p. 275 - 294 Intro to Stoichiometry • Reaction Stoichiometry: Involves the mass relationships between reactants and products in a chemical reaction • Coefficients in a chemical reaction represent the mole ratios

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of each substance that react together • Example: $4\text{Fe} + 3\text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3$

Chapter 9 Stoichiometry Notes

9.3 Objectives Describe a method for determining which of two reactants is a limiting reactant. Calculate the amount in moles or mass in grams of a product, given the amounts in moles or masses in grams of two reactants, one of which is in excess.; Distinguish between theoretical yield, actual yield, and percentage yield.; Calculate percentage yield, given the actual yield and

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