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Chapter 19 Acids And Bases

Introduction to
acids \u0026amp;
bases - Chapter
19 ~~Acids and~~
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Introduction

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Chapter 19:

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By: Loreley

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Acids, Bases,

Oxides \u0026amp;

Ionic Equations

- GCE O Level

Chemistry

Lecture Chapter

14 (Acids and

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Salts Acids and
Bases - Reaction
with each other
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Acid-Base

Equilibria and
Buffer Solutions

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3.9 Buffers Acid-
Base Theories

8.1

Brønsted-Lowry
theory of acids
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(Acids and

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Acid Base and
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2 ||

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~~1 Chapter 16 Acid~~

~~Base Equilibria~~

Acid base \u0026amp;

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Base Class 10th

Science chapter

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7 | Equilibrium
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Equilibrium 01 |

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of acids and

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~~Salts L6 | Doubt~~

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~~Solutions |~~

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Buffer Solution,

pH Calculations,

Henderson

Hasselbalch

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Equation

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Problems FSc

~~Chemistry Book1,~~

~~CH 8, LEC 19:~~

~~Buffer Solutions~~

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~~19 | Baking Soda~~

~~and Washing Soda~~

~~in Hindi~~

~~(????? ???)~~

Chapter 19 Acids

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And Bases

An acid is a substance that contains hydrogen and ionizes to produce hydrogen ions in aqueous solution; A base is a substance that contains a hydroxide group and dissociates to produce a

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hydroxide ion in
aqueous
solution.

Chapter 19:
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Section 19.1 –
Acid-Base
Theories Acids
have a sour
taste, change
the color of an

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indicator, can be strong or weak electrolytes in aqueous solution, and react with metals.

Chapter 19 –
Acids, Bases,
and Salts
states that an
acid is a

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hydrogen ion donor, and a base is a hydrogen ion acceptor. It is a more inclusive model of acids and bases.

conjugate acid. is the species produced when a base accepts a hydrogen ion.

conjugate base.

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is the species
produces when a
acid donates a
hydrogen ion.
conjugate acid-
base pair.

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according to
Arrhenius, acids

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are hydrogen-containing compounds that ionise to yield hydrogen ions (H^+) in aqueous solution. bases are compounds that ionise to yield hydroxide ions (OH^-) in aqueous solution.

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consider the
equation for the
ionization of
hydrogen
fluoride in
water. According
to the Brønsted-
Lowry
definition, the

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equation is
written this
way.

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H₂O u0003
H₃O⁺ u0002 Fu0003

What are the
conjugate acid-
base pairs?

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CHAPTER 19 Acids

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and Bases What
You\u2019ll ...

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STUDY. PLAY.

Tastes sour.

Acid. Changes
the color of an
acid-base

indicator (acid,
base, or both)

Both. Can be
strong or weak

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electrolytes in aqueous solution. Acid. In drinks, citrus, pop, etc. Acid. This acid is associated with protein. amino acid. Used in fragrances and flavors.

Chapter 19:

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Arrhenius Acids.
Acids are
hydrogen
containing
compounds,
ionize to yield
hydrogen ions in

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aqueous solutions. Bases are compounds that ionize to yield hydroxide ions in aqueous solutions.

Monoprotic acids. acids that contain 1 ionizable hydrogen such as nitric acid (HNO₃) diprotic

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Brønsted-Lowry
acids and bases
come in pairs. A
conjugate base

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is the remainder of the B-L acid, after it donates its H^+ . A conjugate acid is the substance formed when the B-L base gains a H^+ . Thus, a conjugate acid-base pair is related by the loss or gain of a single

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hydrogen ion.

Chapter 19 Acids, Bases, and Salts

If the salt is from a strong acid and base then the pH is 7, for example KNO_3 . 3. A salt formed between a strong acid and a conjugate of a

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weak base is an

. acid salt,

example NH_4Cl , pH = acidic.

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Acids, Bases,

and Salts -

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Chapter 19

Acids, Bases,
and Salts. Acids
and Bases:

Basics. These
are really just
a specific type
of chemical
compound. No
mysteries here.

1) They MUST be
ionic compounds,

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and. MUST
dissolve in
water, that's
the only way to
be chemically
active. We call
this.
disassociation.

Chapter 19
Acids, Bases,
and Salts
Chemistry (12th
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to Chapter 19 -
Acids, Bases,
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19 - Acids,

Bases, and ...

Chapter 19

"Acids, Bases,

Page 36/87

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and Salts". Use these activities to study the vocabulary and major concepts presented in this chapter. compounds derived from the reaction of a strong base with a weak acid or a strong acid with a weak base or a

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solution of a
weak base and
one of its
salts.

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notes (fill in
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equilibrium

constants

Chapter 19,

Slides 1-4:

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naming Chapter

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Properties of

Bases Chapter

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10-12:

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calculations,
examples

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Equilibrium,
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and Salts ?What
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properties of
acids and bases?
acids taste
sour, will
change the color
of an acid-base
indicator, and
can be strong or
weak
electrolytes in

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Base Theories

8.1

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theory of acids
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Salt || Acid and

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pH Calculations,

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An acid is a substance that contains hydrogen and ionizes to produce hydrogen ions in aqueous solution; A base is a substance that contains a

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hydroxide group
and dissociates
to produce a
hydroxide ion in
aqueous
solution.

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Acid-Base
Theories Acids

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have a sour taste, change the color of an indicator, can be strong or weak electrolytes in aqueous solution, and react with metals.

Chapter 19 –
Acids, Bases,
Page 56/87

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and Salts

states that an acid is a hydrogen ion donor, and a base is a hydrogen ion acceptor. It is a more inclusive model of acids and bases.

conjugate acid.
is the species produced when a

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base accepts a hydrogen ion.
conjugate base.
is the species produces when a acid donates a hydrogen ion.
conjugate acid-base pair.

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according to Arrhenius, acids are hydrogen-containing compounds that ionise to yield hydrogen ions (H^+) in aqueous solution. bases are compounds that ionise to yield hydroxide ions (OH^-) in

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aqueous
solution.

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and salts

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equation for the
ionization of
hydrogen
fluoride in
water. According

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to the Brønsted-Lowry definition, the equation is written this way.

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H₂O u0003
H₃O⁺ u0002 F⁻ u0003

What are the conjugate acid-base pairs?

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You\u2019ll ...

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and Salts.

STUDY. PLAY.

Tastes sour.

Acid. Changes
the color of an
acid-base

indicator (acid,

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base, or both)

Both. Can be strong or weak electrolytes in aqueous

solution. Acid.

In drinks,

citrus, pop,

etc. Acid. This

acid is

associated with

protein. amino

acid. Used in

fragrances and

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Acids are
hydrogen
containing

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compounds,
ionize to yield
hydrogen ions in
aqueous
solutions. Bases
are compounds
that ionize to
yield hydroxide
ions in aqueous
solutions.

Monoprotic
acids. acids
that contain 1
ionizable

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hydrogen such as
nitric acid
(HNO_3) diprotic
acids.

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related by the .
loss or gain. of
a single
hydrogen ion.

Chapter 19 Acids, Bases, and Salts

If the salt is
from a strong
acid and base
then the pH is
7, for example
KNO₃. 3. A salt

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formed between a strong acid and a conjugate of a weak base is an . acid salt,

example NH_4Cl , pH =acidic.

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and Salts. Acids
and Bases:

Basics. These
are really just
a specific type
of chemical
compound. No

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mysteries here.

1) They MUST be ionic compounds, and. MUST dissolve in water, that's the only way to be chemically active. We call this. disassociation.

Chapter 19
Acids, Bases,
Page 77/87

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and Salts". Use
these activities
to study the
vocabulary and
major concepts
presented in
this chapter.
compounds
derived from the
reaction of a
strong base with

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a weak acid or a strong acid with a weak base or a solution of a weak base and one of its salts.

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