

Chapter 11 Review Molecular Composition Of Gases Section 11 2 Answers

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334 CHAPTER 11 FIGURE 11-1 At the same temperature and pressure, balloons of equal volume have equal numbers of molecules, regardless of which gas they contain. Hydrogen molecule 1 mol H 2 at STP = 22.4 L Oxygen molecule 1 mol O 2 at STP = 22.4 L Carbon dioxide molecule 1 mol CO 2 at STP = 22.4 L

CHAPTER 11 Molecular Composition of Gases

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Chapter 11 - Molecular Composition of Gases 11-1 Volume-Mass Relationships of Gases I. Measuring and Comparing the Volumes of Reacting Gases A. Observations of Gay-Lussac 1. 2 liters H 2 + 1 liter O 2 \u00c2 liters H 2O vapor 2. 2 volumes H 2 + 1 volume O 2 \u00c2 2 volumes H 2O vapor 3. 1 volume H 2 + 1 volume Cl 2 \u00c2 2 volumes HCl 4.

Chapter 11 - Molecular Composition of Gases

CHAPTER 11 REVIEW Molecular Composition of Gases MIXED REVIEW SHORT ANSWER Answer the following questions in the space provided 1 c The average speed of a gas molecule is most directly related to the (a) polarity of the molecule (b) pressure of the gas (c) temperature of the gas (d) Chapter 11 Review Gases Mixed Answer Key

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Chapter 11 Review Molecular Composition Of Gases Mixed Answers Answer. Molar mass = 151.1 g/mol Potassium: (39.10 / 101.1) x 100 = 38.67% Nitrogen: (14.01 / 101.1) x 100 = 13.86% Oxygen: (48.00 / 101.1) x 100 = 12.58%.

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