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pH = 14.000 -

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 $\text{pOH} = 14.000 - 2.187 = 11.813$  4)

-A solution is created by measuring  $3.60 \times 10^{-3}$  moles of NaOH and  $5.95 \times 10^{-4}$  moles of HCl into a container and then water is added until the final volume is 1.00 L. Calculating pH and pOH

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pOH = 14.00 -  
4.21 = 9.79 6) +A

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 $= -\log(6.2 \times 10^{-5}) =$

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solution with a H  
concentration of  
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$$\text{pOH} = 14.000 -$$

$$2.187 = 11.813 \text{ 4)}$$

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NaOH and  $5.95 \times$

$10^{-4}$  moles of HCl

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