

## Butler Ionic Equilibrium Solubility Chapter 6

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At a certain temperature, the solubility of  $\text{Fe}(\text{OH})_2$  in water is  $7.7 \times 10^{-6} \text{ mol/L}$  (M). Its  $K_{sp}$  can be calculated based on the

equilibrium equation:  $\text{Fe}(\text{OH})_2 \rightleftharpoons \text{Fe}^{2+} + 2\text{OH}^-$   $\text{Fe}(\text{OH})_2 \rightleftharpoons \text{Fe}^{2+} + 2\text{OH}^-$ . Therefore, the solubility product expression is:  $K_{sp} = [\text{Fe}^{2+}][\text{OH}^-]^2$   $K_{sp} = [\text{Fe}^{2+}][\text{OH}^-]^2$ .

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